**Effective vs non-effective** writing

![solid-liquid-gas-picture[1]]()

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| **Text A** |
| All matter may be classified as either a solid, a liquid or a gas. The atoms or molecules of a solid are densely packed and have very little freedom of movement. Most solids require a considerable force in order to change the shape or volume of most of them. Rubber, wood, glass, iron, cotton, and sand are firm and have a definite form and so are all classified as solids. |
| Air, oxygen, and carbon dioxide are gases and have no fixed shape or volume of their own. Any container can be filled by these as they diffuse and spread out to fill them. The atoms or molecules of gases move very rapidly and are widely spaced. If water is put into a tyre, it will run to the bottom; if a tyre is filled with air, it fills the whole space inside it. They therefore expand or compress to fit any area. |
| Crystalline solids and amorphous solids are the two classes solids may be further divided into. When heated, the change to a liquid known as melting is sharp and clear for crystalline solids. Rocks, wood, paper and cotton are this type. These are made up of atoms arranged in a definite pattern. Rubber, glass and sulphur are amorphous solids. These gradually soften when heated, as the pattern of the atoms is not orderly. |
| A definite form does not characterise liquids. If a table has water, milk, or oil poured on it, these would flow all over the surface. So it will take the shape of any container in which it is poured since these atoms are loosely structured and they do not keep their shape. The atoms or molecules attract each other, meaning they are able to flow. |

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| **Text B** |
| All matter may be classified as either a solid, a liquid or a gas. Solids are firm and have a rigid form. Rubber, wood, glass, iron, cotton, and sand are all classified as solids. The atoms or molecules of a solid are densely packed and have very little freedom of movement: therefore, most solids require a considerable force in order to change their form or volume. |
| Solids may be further divided into two classes: crystalline and amorphous. Crystalline solids include rocks, wood, paper and cotton. These solids are made up of atoms arranged in a definite pattern. When crystalline solids are heated, the change to liquid, known as melting, is sharp and clear. Amorphous solids include rubber, glass and sulphur. The pattern of the atoms in amorphous solids is not orderly: as a result, when they are heated they soften gradually. |
| Liquids, on the other hand, do not have rigid form. If water, milk, or oil is poured on a table, it will flow all over the surface. The atoms or molecules of liquids attract each other and thereby enable liquids to flow. However, the atoms in liquids are loosely structured and they do not keep their shape. Therefore, a liquid will take the shape of any container in which it is poured. |
| Gases, such as air, oxygen, and carbon dioxide, are unlike both solids and liquids because they do not have a rigid form or fixed volume of their own. Volumes vary because gases diffuse as they spread out to fill any container. For example, if water (liquid) is put into a tyre, it will run to the bottom; if air (gas) is put into a tyre, it fills the whole space inside. The atoms or molecules of gases are widely spaced and move very rapidly. These atoms therefore expand or compress to fit any area |

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